Procedure Questions

1. Small drops of a liquid has formed around the top of the test tube, the cobalt paper determines that this liquid is water.
2. In the bottom of the test tube is a pink and blue solid.

5. A cloudy substance is formed
6. Barium Sulphate is created
7. The wire turns red in the fire, and goes black when it is removed
8. Oxygen in the air
9. Bubbles are produced
10. A loud popping sound
11. Hydrogen
12. Bubbles are produced
13. The split re ignites
14. Oxygen
15. Silver begins to form on the copper foil
16. Water slowly condenses on the edges
17. The limewater goes cloudy

Discussion Questions

1. Cobalt (II) Chloride dihydrate 🡪 Cobalt(II) Chloride + Water

Sodium Sulphate + Barium Chloride 🡪 Sodium Chloride + Barium Sulphate

Copper(I) + Oxygen 🡪 Copper(I) Oxide

Hydrochloric Acid + Zinc 🡪 Zinc Chloride + Hydrogen

Hydrogen Peroxide 🡪 Hydrogen + Oxygen

Silver Nitrate + Copper 🡪 Copper(II) Nitrate + Silver

1. CoCl2 (s) • 2H2O(l) 🡪 CoCl2 (s)+ 2H2O(g)

2Na2SO4 (aq)+ 2BaCl2 (aq)🡪 4NaCl (aq) + 2BaSO4 (s)

4Cu (s) + O2 (g) 🡪 2Cu2O (s)

2HCl (aq) + 2Zn (s) 🡪 2ZnCl (l) + H2 (g)

H2O2 (l) + H2 (g) 🡪 H2 (g) + O2 (g)

Ag(NO3)2 (l) + Cu (s) 🡪 Cu(NO3)2 (l) + Ag (s)